

VIDEO PEMBELAJARAN
MATERI: CHEMICAL KINETIC

PART 2

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Jaslin Ikhsan: Chemical Kinetic Part 2

Rate laws can be complicated

1) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightarrow 2\text{HI}(\text{g}) \quad R = k[\text{H}_2][\text{I}_2]$

2) $\text{H}_2(\text{g}) + \text{Br}_2(\text{g}) \rightarrow 2\text{HBr}(\text{g}) \quad R = \frac{2k[\text{H}_2][\text{Br}_2]^{1/2}}{1 + k'[\text{HBr}][\text{Br}_2]^1}$

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Finding rate laws experimentally

There are two common methods for determining rate laws:

Method of isolation
Set up reaction so one reactant is in excess. Any change in rate will be due to changes in other reactant. Repeat for other reactant.

$$R = k'[B]^n \quad \text{where} \quad k' = k[A]^m$$

Method of initial rates
Measure concentration change as a function of time, $\gamma(t)$, for a series of experimental conditions. (Conditions must include sets where the reactant A has the same initial concentration but B changes and vice versa).

02:00 / 15:04 360p

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